

GCSE

Chemistry B (Twenty First Century)

Unit **J258/03**: Higher Tier – Breadth in chemistry

General Certificate of Secondary Education

Mark Scheme for June 2018

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This mark scheme is published as an aid to teachers and students, to indicate the requirements of the examination. It shows the basis on which marks were awarded by examiners. It does not indicate the details of the discussions which took place at an examiners' meeting before marking commenced.

All examiners are instructed that alternative correct answers and unexpected approaches in candidates' scripts must be given marks that fairly reflect the relevant knowledge and skills demonstrated.

Mark schemes should be read in conjunction with the published question papers and the report on the examination.

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Annotations available in RM Assessor

Annotation	Meaning
	Correct response
	Incorrect response
	Omission mark
	Benefit of doubt given
	Contradiction
	Rounding error
	Error in number of significant figures
	Error carried forward
	Level 1
	Level 2
	Level 3
	Benefit of doubt not given
	Noted but no credit given
	Ignore

Abbreviations, annotations and conventions used in the detailed Mark Scheme (to include abbreviations and subject-specific conventions).

Annotation	Meaning
/	alternative and acceptable answers for the same marking point
✓	Separates marking points
DO NOT ALLOW	Answers which are not worthy of credit
IGNORE	Statements which are irrelevant
ALLOW	Answers that can be accepted
()	Words which are not essential to gain credit
—	Underlined words must be present in answer to score a mark
ECF	Error carried forward
AW	Alternative wording
ORA	Or reverse argument

Subject-specific Marking Instructions**INTRODUCTION**

Your first task as an Examiner is to become thoroughly familiar with the material on which the examination depends. This material includes:

- the specification, especially the assessment objectives
- the question paper
- the mark scheme.

You should ensure that you have copies of these materials.

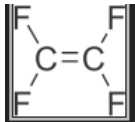
You should ensure also that you are familiar with the administrative procedures related to the marking process. These are set out in the OCR booklet **Instructions for Examiners**. If you are examining for the first time, please read carefully **Appendix 5 Introduction to Script Marking: Notes for New Examiners**.

Please ask for help or guidance whenever you need it. Your first point of contact is your Team Leader.

The breakdown of Assessment Objectives for GCSE (9-1) in Chemistry B:

	Assessment Objective
AO1	Demonstrate knowledge and understanding of scientific ideas and scientific techniques and procedures.
AO1.1	Demonstrate knowledge and understanding of scientific ideas.
AO1.2	Demonstrate knowledge and understanding of scientific techniques and procedures.
AO2	Apply knowledge and understanding of scientific ideas and scientific enquiry, techniques and procedures.
AO2.1	Apply knowledge and understanding of scientific ideas.
AO2.2	Apply knowledge and understanding of scientific enquiry, techniques and procedures.
AO3	Analyse information and ideas to interpret and evaluate, make judgements and draw conclusions and develop and improve experimental procedures.
AO3.1	Analyse information and ideas to interpret and evaluate.
AO3.1a	Analyse information and ideas to interpret.
AO3.1b	Analyse information and ideas to evaluate.
AO3.2	Analyse information and ideas to make judgements and draw conclusions.
AO3.2a	Analyse information and ideas to make judgements.
AO3.2b	Analyse information and ideas to draw conclusions.
AO3.3	Analyse information and ideas to develop and improve experimental procedures.
AO3.3a	Analyse information and ideas to develop experimental procedures.
AO3.3b	Analyse information and ideas to improve experimental procedures.

Question			Answer	Marks	AO element	Guidance
2	(a)		No overall loss (in mass) idea / No elements/mass/atoms/chemicals/particles/compounds lost / law states that matter is neither (created nor) destroyed in a chemical reaction / AW ✓ Carbon dioxide is a gas / Carbon dioxide leaves the test tube / a gas is given off / idea that all products are not in the test tube / AW ✓	2	3.1b × 2	ALLOW It is an open system
	(b)		FIRST CHECK ANSWER ON ANSWER LINE If answer = 52.2 / 52.4 / 52.3 (%) award 4 marks (formula mass of reactants or MgCO ₃) = 84.3/84 ✓ (formula mass of product or CO ₂) = 44 ✓ Correct substitution = 44/84.3 x 100 / 44/84 x 100 ✓ Ans+dec pl= 52.2 / 52.4 / 52.3 (%) (1 decimal place) ✓	4	2.2 × 3 <	

Question		Answer	Marks	AO element	Guidance
3	(a)	A high breaking strength <u>is good</u> / is strong(er) / won't break / cup would not hold its shape / can hold boiling water / softens <u>above 100</u>/at higher temperature than B or C ✓ B low breaking strength <u>is not good</u> / weak(er) / breaks / would not hold its shape / could not hold boiling water /softens <u>below 100</u>/at lower temperature than A or C ✓ C low breaking strength <u>is not good</u> / weak(er) / breaks / would hold its shape / could hold boiling water / softens <u>above 100</u>/above B/below A ✓ Breaking strength	3	3 × 3.1a	IGNORE list of properties repeated from the table e.g. 'A has high breaking strength, is fairly flexible but softens at 250' = 0 Answer must show some processing of information e.g. links properties to 'good' and 'bad' or interprets properties For (3) marks answer must discuss at least two properties
	(b)	B ✓ Lowest softening temperature ✓	2	2 × 3.2a	IGNORE references to flexibility/strength
	(c)		1	2.1	ALLOW f

Question			Answer	Marks	AO element	Guidance
4	(a)	(i)	No/very little/<1% oxygen until 2.5 billion years ago/for about (first) 1.5 billion years AW ✓ Fluctuations but overall increase / up and down but overall rise ✓	2	2 × 3.1a	ALLOW any number in range 2.8 – 2.5 for ‘about 2.5 / 1.2-1.5 for ‘about 1.5’ IGNORE ‘increases and decreases’ or ‘up and down’ alone
		(ii)	1.7 – 2(.0)	1	3.1a	
		(iii)	Plants/bacteria evolved / photosynthesis✓ plants/bacteria established / number of plants or bacteria stayed the same / animals evolved / animals use up oxygen / respiration AW ✓	2	2 × 1.1	
	(b)		$2\text{FeS}_2 + 7\text{O}_2 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{SO}_2$	1	1.2	

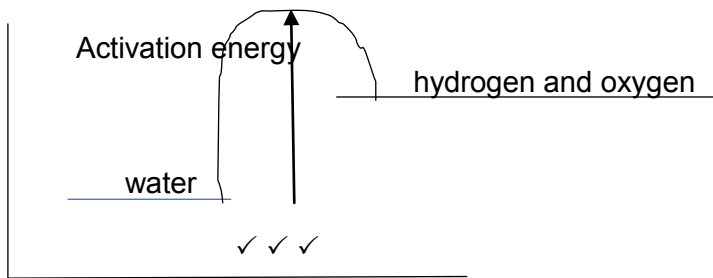
Question			Answer	Marks	AO element	Guidance
5	(a)		Bromine/Br₂ identified ✓ displaced (by the chlorine) / chlorine is more reactive than bromine ✓ $Cl_2 + 2Br^- \rightarrow Br_2 + 2Cl^-$ ✓	3	3 × 1.1	ALLOW from equation IGNORE 'replace' DO NOT ALLOW '<u>chloride</u> displaces bromide' ALLOW 'chlorine displaces bromide' for MP2 IGNORE K⁺ ions
	(b)		Astatine reacts with sodium to form NaAt ✓ Astatine is less reactive than iodine ✓	1	2.1	

Question			Answer	Marks	AO element	Guidance
6	(a)		large surface area to volume ratio / (very) small but have a large surface area ✓	1	1.1	
	(b)		For: (More complete combustion hence) less pollutants / less harmful gases / less incomplete combustion / less named pollutant: (carbon) particulates / carbon monoxide / CO, unburnt fuel/hydrocarbons AW ✓ Carbon monoxide is toxic//blocks haemoglobin / CO or particulates or unburnt HCs cause breathing or respiratory difficulties / particulates cause asthma/breathing difficulties / may cause cancer etc ✓ Against: CeO₂/nanoparticles (may be) harmful / toxic / risks not known ✓	3	3.1b 1.1 3.1b	IGNORE 'more complete combustion' alone (repeats Q) IGNORE 'less pollution' ALLOW idea of nanoparticles may harm humans, plants, animals or the environment IGNORE 'nanoparticles may be pollutants /cause pollution' (too vague)
	(c)		Charges balance/neutral/cancel / oxygen gives an overall charge of 4- / there are two O ²⁻ ions AW ✓	1	1.2	IGNORE charges are equal
	(d)		FIRST CHECK ANSWER ON ANSWER LINE If answer = 3×10^{-21} (mol) award 3 marks Molecules in one nanoparticle = $8 \times 10^{-27} / 4 \times 10^{-30} / = 2000$ ✓ Recall Avogadro constant = $6(.0) \times 10^{23}$ ✓ Moles = $2 \times 10^3 / 6 \times 10^{23} = 3.3 \times 10^{-21}$ (mol) ✓	3	2.2 1.2 2.2	ALLOW full credit (3) marks for any answer which starts by cubing volumes ALLOW correct working shown OR 2000 for (1) ALLOW ECF

Question			Answer	Marks	AO element	Guidance
7	(a)	(i)	Equilibrium sign / $\rightleftharpoons$ / arrows point both ways ✓	1	1.1	ALLOW answers referring to the sign/arrow IGNORE it is reversible
		(ii)	forward and back(ward) reactions / both directions / reactants(N_2 and O_2) forming products (NO) <u>and</u> products forming reactants ✓ rates are equal ✓	2	2 × 1.1	ALLOW equations for the correct reactions
	(b)	(i)	Fertilisers / explosives	1	1.1	ALLOW for growth of/nitrates for plants but IGNORE plants alone/ammonia/Haber process
		(ii)	No change ✓ Equal moles/molecules/particles on each side (of the equation) ✓	2	2 × 1.1	Mark separately
	(c)		Any two FOR: 100% atom economy / all reactants used up idea ✓ No by-products / no waste ✓ raw materials come from the air ✓ sustainable ✓ works at low pressure / AW ✓ Any one AGAINST: (very) high temperature (needed) / needs a lot of energy/fuel / doesn't give 100% yield / low yield ✓	3	3 × 2.1	ALLOW 'high atom economy' IGNORE 'readily available' IGNORE 'renewable'

Question			Answer	Marks	AO element	Guidance
8	(a)	(i)	$\text{MnO}_2(\text{s}) + 2\text{C}(\text{s}) \rightarrow 2\text{CO}(\text{g}) + \text{Mn}(\text{s})$ correct formulae and balancing ✓ state symbols ✓	2	2 × 1.2	ALLOW state symbol mark for any version of manganese oxide + carbon → carbon oxide + manganese
		(ii)	Manganese is less reactive than carbon ORA ✓ carbon reduces / removes oxygen from / donates electrons to manganese (oxide) ✓	2	2 × 2.1	
	(b)		strong forces/bonds/attraction / electrostatic attractions between ✓ (free/delocalised/sea of) electrons ✓ and positive ions (from metal) ✓	3	3 × 1.1	DO NOT ALLOW intermolecular forces IGNORE metal atoms DO NOT ALLOW protons/nuclei

Question			Answer	Marks	AO element	Guidance
9	(a)		(contains) potassium / K ✓	1	1.2	
	(b)	(i)	lines ✓ (lines are) coloured / (lines are) on a dark/black background ✓	2	2 × 1.1	ALLOW 'series/range of colours' IGNORE 'on a white background'
		(ii)	Compare/match the spectrum with the known spectrum (of potassium/(ECF answer to (a)) / AW ✓	1	1.2	
	(c)	(i)	White precipitate	1	1.2	IGNORE cream
		(ii)	$\text{BaCl}_2 + \text{Na}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$ correct formula for one product $\text{BaSO}_4/\text{NaCl}$ ✓ fully correct equation with balancing ✓	2	2 × 1.2	

Question			Answer	Marks	AO element	Guidance
10	(a)		Hydrogen is produced at the cathode Water contains H^+ and OH^- ions	1	1.1	Both needed
	(b)		FIRST CHECK ANSWER ON ANSWER LINE If answer = (+)490 (kJ/mol) award 3 marks (bonds broken): (= 4(O–H)) OR 1856 (kJ/mol) ✓ (bonds made): (= 2(H–H) + O=O) OR 1366 (kJ/mol) ✓ energy change (= 1856 – 1366) / (bonds broken – bonds made) = (+) 490 (kJ/mol) ✓	3	3 × 2.2	For MP1 and MP2 ALLOW either correct working or correct value DO NOT ALLOW –490 (if sign is shown, must be correct) ALLOW ECF for maximum (2) If answer should be negative (by ecf), then negative sign must be shown for third mark to be scored. Plus sign is not necessary.
	(c)			3	2 × 2.2 1.1	Water on LHS lower than hydrogen and oxygen higher on RHS ✓ ALLOW ECF for exothermic diagram from a negative answer to (b) DO NOT PENALISE correct diagram if answer to (b) is incorrect Curve with hump ✓ Activation energy labelled with arrow starting at reactants and ending at peak of hump ✓ ALLOW double arrow / line with end stops DO NOT ALLOW downwards arrow

Question			Answer	Marks	AO element	Guidance
11	(a)		$\text{Al}^{3+} + 3\text{e} \rightarrow \text{Al} \checkmark$ $2\text{O}^{2-} \rightarrow \text{O}_2 + 4\text{e} \checkmark$	2	2 × 1.2	ALLOW equations with electrons on the right (eg $\text{Al}^{3+} \rightarrow \text{Al} - 3\text{e}$) ALLOW $\text{O}^{2-} \rightarrow \frac{1}{2}\text{O}_2 + 2\text{e}$ ALLOW (1) mark if number of electrons are correct for both equations but on incorrect sides of both equations (even if oxygen is shown as O)
	(b)		ions / charged particles $\checkmark$ can't move in solid / held in lattice / do not move / in fixed positions $\checkmark$ can move in liquid/when molten $\checkmark$	3	3 × 1.1	DO NOT ALLOW electrons/'particles' alone
	(c)		FIRST CHECK ANSWER ON ANSWER LINE If answer = 667etc (dm³) award 4 marks $4 \times 27(\text{g}) (=108) \checkmark$ $=1000 \div 108 \text{ ECF } (= 9.259) \checkmark$ $3 \times 24 (\text{dm}^3) (=72) \checkmark$ $= (\times 72 \text{ ECF } =) 667 (\text{dm}^3) \checkmark$ OR No moles Al = $1000/27 (=37.04) \checkmark$ Uses ratio 4:3 in calculation $\checkmark$ No moles O ₂ = $\text{ANS} \times \frac{3}{4} (=27.77\text{.....}) \checkmark$ $(\times 24=) 667 (\text{dm}^3) \checkmark$	4	4 × 2.2	$9.259 = (2)$ for MP1 and MP2 $0.667 = (3)$ for MP1 MP3 and MP4 $222 = (3)$ for MP1 MP2 and MP4 $0.222 = (2)$ for MP1 and MP4 ALLOW 666 (Rounding already assessed in earlier Q) ALLOW any number of sig figs $27.7 = (3)$ for MP1, MP2 and MP3

			Answer	Marks	AO element	Guidance
12	(a)		(The arrow shows that) the acid is fully dissociated/ionised / the reaction goes to completion AW / there is no equilibrium sign ✓	1	1.1	
	(b)	(i)	Volumetric pipette ✓	1	1.2	
		(ii)	FIRST CHECK ANSWER ON ANSWER LINE If answer = 0.0025 (mol) award 3 marks Unit conversion cm^3 to $\text{dm}^3 = 25 \times 10^{-3} = 0.025 \text{ dm}^3$ / divides 25 by 1000 / shows 25×0.001 in working ✓ Shows $\times 0.1$ in working ✓ Correct answer = 0.0025 (mol) ✓	3	1.2 2 × 2.2	ALLOW ECF
		(iii)	FIRST CHECK ANSWER ON ANSWER LINE If answer = 0.051 (mol / dm³) award 3 marks moles $\text{H}_2\text{SO}_4 = \text{ANS (ii)} / 2 / 0.0025 / 2 (= 0.00125)$ ✓ conc $\text{H}_2\text{SO}_4 = 0.00125$ (ECF) / 0.0245 ✓ conc $\text{H}_2\text{SO}_4 = 0.051$ (mol / dm ³) ✓	3	3 × 2.2	ALLOW ECF (including from part (ii)) 0.102 = (1) for MP2 ALLOW answer with working to 2 sig figures

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